

Lewis Dot Practice  
CHM 1045

Draw the Lewis dot structures for the following including

- a. Bonds
- b. Lone pairs
- c. Formal charges

1.  $\text{PBr}_3$
2.  $\text{N}_2\text{H}_2$
3.  $\text{CH}_3\text{OH}$
4.  $\text{NO}_2^-$
5.  $\text{C}_2\text{H}_4$
6.  $\text{C}_2\text{H}_2$
7.  $\text{ONCl}$
8.  $\text{BrO}_3^-$
9.  $\text{HBr}$
10.  $\text{N}_2$
11.  $\text{C}_2\text{H}_5\text{OH}$
12.  $\text{N}_2\text{F}_4$
13.  $\text{SF}_6$
14.  $\text{SCl}_2$
15.  $\text{ICl}_3$
16.  $\text{Br}_2$
17.  $\text{CO}_2$
18.  $\text{O}_2$
19.  $\text{PCl}_5$
20.  $\text{SO}_4^{2-}$

